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Chemistry Unit 8  
Worksheet 3: Adjusting to Reality - Limiting Reactant 1. Write the balanced equation for the reaction between hydrogen and oxygen.  
Balanced Equation:  $2 +$

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1 2 H<sub>2</sub>O(g) Suppose  
that 4 molecules of  
hydrogen gas and 4  
molecules of oxygen  
gas react to form  
water. Make a drawing  
that represents the  
reaction container  
before and after the

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Chemistry Unit 8

Worksheet 4 Samples

of Every Kind of

Problem On a separate

sheet of paper, write a

complete solution to

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each of the problems  
below. Follow the  
procedure outlined in  
class. ...  $13.3 \text{ g} \times 1$   
 $\text{mole} = 0.416 \text{ mole O}_2$   
 $32 \text{ g} \times 0.277 \text{ mole} \times 81.4$   
 $\text{g} = 22.5 \text{ g ZnO}$  1 mole  
 $0.200 \text{ mole} \times 97.5 \text{ g} =$   
 $19.5 \text{ g ZnS}$  1 mole  
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## **Unit 8 Worksheet 4 - Studylib**

In the early 1990s,  
after a decade of  
education research to

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develop and validate

Instruction(TM),

physicist David

Hestenes was awarded

grants from the

National Science

Foundation for another

decade to spread the

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program nationwide.

As of 2019,

approximately 14,000

teachers have

participated in summer

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development involving  
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8 - Worksheet 5  
Samples of Every Kind  
of Problem On a  
separate sheet of  
paper, write a



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complete solution to  
each of the problems  
below. Follow the Date  
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Chemistry I Name

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WS 3 Yet More

Balancing Equations

Write balanced

chemical equations for  
the following reactions.

1. Ammonia ( $\text{NH}_3$ )

reacts with aqueous

hydrogen chloride to

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form ammonium

chloride.  $\text{NH}_3 + \text{HCl} \rightarrow$

$\text{NH}_4\text{Cl}$

Calcium

carbonate decomposes

upon heating to form

calcium oxide and

carbon dioxide.  $\text{CaCO}_3$

$\rightarrow \text{CaO} + \text{CO}_2$

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**Chemistry I**

**Name\_KEY Unit 7 WS**

**3 Yet More ...**

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Worksheet 2 -



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Measuring Pressure .

Problem 1 and 2

Calculate the pressure using the barometer .

1.  $720 \text{ mmHg} = 0.95 \text{ atm}$ . 2.  $760 \text{ mmHg} = 1 \text{ atm}$ . Problems 3 and 4.

Calculate the pressure of the gas in the flask connected to the manometer. 3. a.

$127 - 84 = 43 \text{ mmHg}$   
 $730 + 43 = 773 \text{ mmHg}$

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v2.0 10. Barium nitrate reacts with aqueous sodium sulfate to produce solid barium sulfate and aqueous sodium nitrate. Abigail places 20.00 mL of 0.500 M barium nitrate in a flask. She has a 0.225M sodium sulfate solution available.

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U8 ws 3 v2.0 10.

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